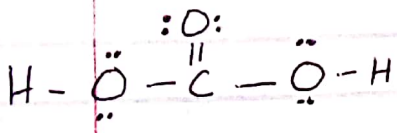
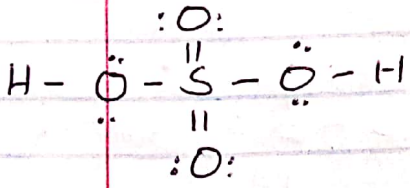
Page 1 (Q1 - Q3)

Draw the lewis dot structure for the following:

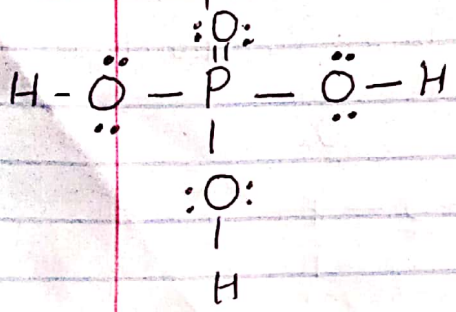
1 Carbonic acid.



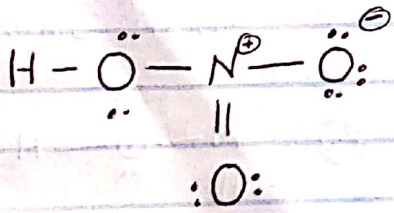
(ii). Sulphuric acid.



(iii) Phosphoric acid.

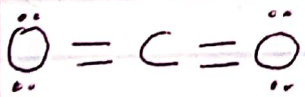


(iv) Nitric acid.

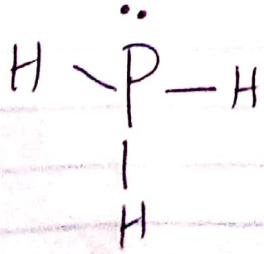


(2) Dot - Structure of Molecules.

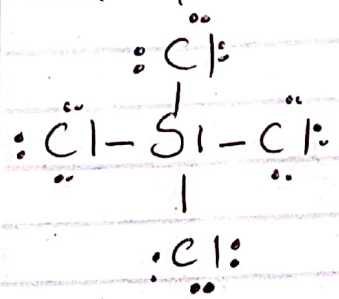
a) $CO_{2(g)}$



(b) PH_3.

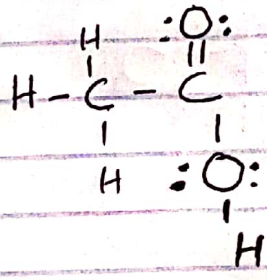


$SiCl_4$.

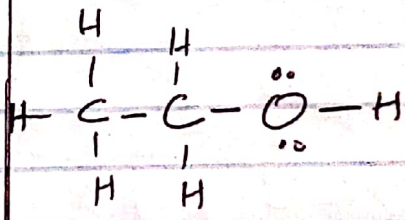


(3) lewis dot structure of the following

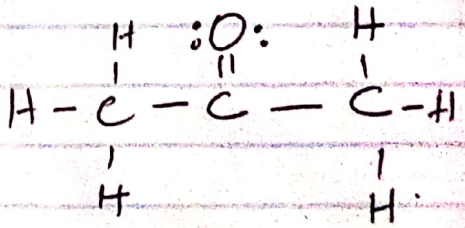
Acetic acid



(b) Ethanol



(c) Acetone.

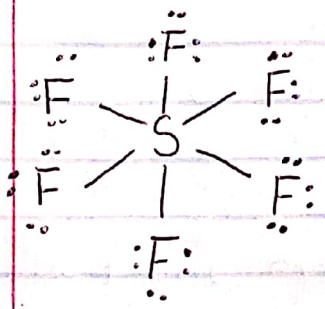


(4)

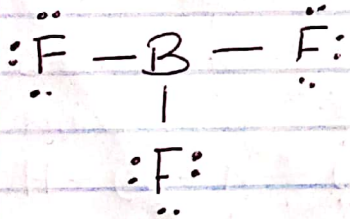
Scanned with CamScanner

4). Draw the dot lewis structure of the following.

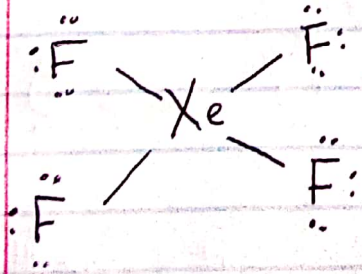
a) SF_6



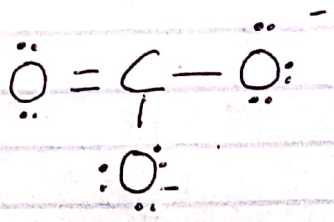
b) BF_3



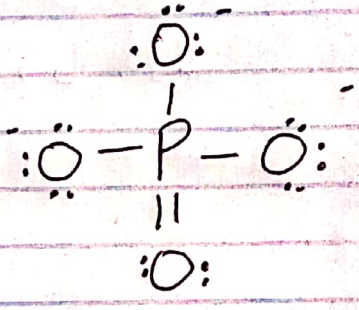
c) XeF_4



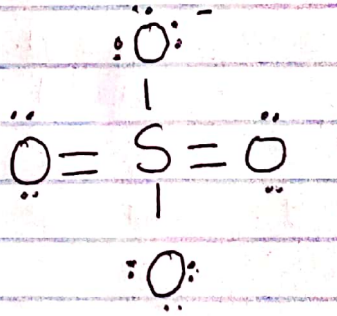
5) CO_3^{2-}.



Phosphate PO_4^{3-}



Sulphate SO_4^{2-}



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Q6

Bond breaking is endothermic $\Delta H = +$

Bond formation is exothermic $\Delta H = -$

EQuation:

$$3H_{2(g)} + N_{2(g)} \longrightarrow 2NH_{3(g)}$$

Bond broken. are:

$3(H-H)$ $= 436.4 \times 3 = 1309.2$

$N \equiv N$ $= 941.4 \times 1 = 941.4$

$+ 2250.6\ KJmol^{-1}$

Bond formation:

$3(N-H) \times 2$ $= 393 \times 3 \times 2 = -2358\ KJ$

The $\Delta H_f = 2250.6 + -2358 =$

$= -107.4\ KJmol^{-1}$

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