

(2) 9.624×10^{-13}

(4) $1.4283 \times 10^{-9} \text{ m}$

(5) ~~1.837~~

~~1.837~~ $1.837 \times 10^{-19} \text{ J}$

6 (a) -0.544 eV

(b) $1.483 \times 10^{-34} \text{ J}\cdot\text{s}$

(c) $5.932 \times 10^{-34} \text{ J}\cdot\text{s}$

8 $l = 6$

$n = -3$

9 $n = 5$ can hold 46 electrons

12 $1.36 \times 10^{-12} \text{ m}$

13 $z = 12$

Carbon

14 1.957×10^{17} Photons